Chemistry 141 Name

Dr. Cary Willard

Quiz 5A (20 points) March 13, 2013

Data: PV = nRT, R = 0.0821 L atm/mol K = 62.4 L torr/mol K, K=oC + 273.16, $\frac{rate 1}{rate 2}=\sqrt{\frac{M\_{2}}{M\_{1}}}$

1. (5 points) An experiment shows that a 364 mL gas sample has a mass of 0.747 g at a pressure of 0.644 atm and a temperature of 37oC. What is the molar mass of the gas?
2. (5 points) A 485 mL flask contains pure neon at a pressure of 539 torr. A second flask with a volume of 395 mL contains pure xenon at a pressure of 627 torr. If the two flasks are connected through a stopcock and the stopcock is opened, what are the partial pressures of each gas and the total pressure?
3. (5 points) A sample of CO2 effuses from a container in 67 seconds. How long would it take the same amount of butane gas (C4H10) to effuse from the same container under identical conditions?
4. (5 points) Which postulate of kinetic molecular theory breaks down under conditions of low temperature? Explain.

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Quiz 5B (20 points) March 13, 2013

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1. (5 points) An experiment shows that a 277 mL gas sample has a mass of 0.615 g at a pressure of 0.504 atm and a temperature of 37oC. What is the molar mass of the gas?
2. (5 points) A 485 mL flask contains pure neon at a pressure of 486 torr. A second flask with a volume of 395 mL contains pure xenon at a pressure of 731 torr. If the two flasks are connected through a stopcock and the stopcock is opened, what are the partial pressures of each gas and the total pressure?
3. (5 points) A sample of CO2 effuses from a container in 82 seconds. How long would it take the same amount of butane gas (C4H10) to effuse from the same container under identical conditions?
4. (5 points) Which postulate of kinetic molecular theory breaks down under conditions of low temperature? Explain.